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| http://ths.sps.lane.edu/chemweb/chemlogo2.jpg | http://ths.sps.lane.edu/chemweb/Atom-02.gif | **Unit 1 Significant Figures Quiz source:** [**http://ths.sps.lane.edu/chemweb/unit1/problems/significantfigures/**](http://ths.sps.lane.edu/chemweb/unit1/problems/significantfigures/) |

Multiple Choice (Choose the best answer.)

1. The measurement, 206 cm, has how many significant (measured) digits?

 1

 2

 3

 4

 5

1. The measurement, 206.0 oC, has how many significant digits?

 1

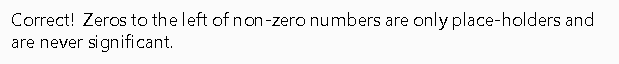
 2

 3

 4

 5

1. The measurement, 0.00206 g, has how many significant digits?

 1 

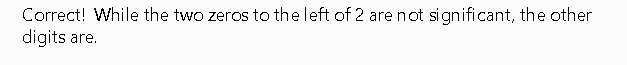
 2

 3

 4

 5

1. The measurement, 0.0020600 mole, has how many significant digits?

 1 

 2

 3

 4

 5

1. The measurement, 2.060 x 10-3coulombs, has how many significant digits?

 1 

 2

 3

 4

 5

1. The measurement, 20600 molecules, has how many significant digits?

 1

 2

 3

 4

 5

1. Add the following three numbers and report your answer using significant figures:

2.5 cm + 0.50 cm + 0.055 cm = ?

 3.055 cm

 3.06 cm

 3.1 cm

 3.0 cm

 3 cm

1. Subtract the following numbers and report your answer using significant figures:

416 g - 210 g = ?

 206. g

 206 g

 210. g

 210 g

 200 g

1. Multiply the following three numbers and report your answer to the correct number of significant figures:

0.020 cm x 50 cm x 11.1 cm = ?

 10 cm3

 11 cm3

 11. cm3

 11.1 cm3

 11.10 cm3

1. Divide the following three numbers and report your answer to the correct number of significant figures:

0.530 g / 0.1010 mL = ?

 2 g/mL

 5.2 g/mL

 5.3 g/mL

 5.25 g/mL

 5.248 g/mL

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| --- | --- | --- | --- |
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